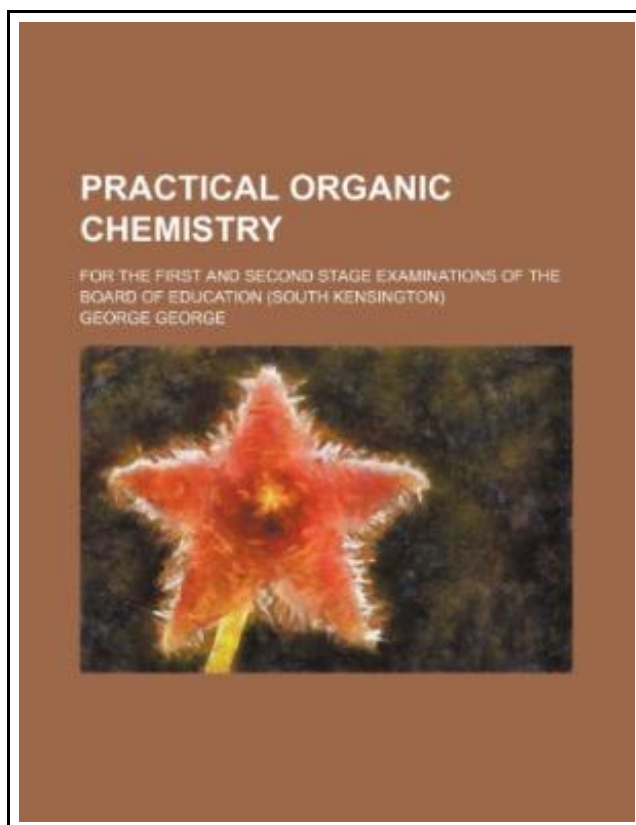


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Reviews


This ebook is very gripping and fascinating. It is among the most awesome ebook i have go through. I found out this publication from my i and dad advised this ebook to understand.
(Olen Shields PhD)


PRACTICAL ORGANIC CHEMISTRY; FOR THE FIRST AND SECOND STAGE EXAMINATIONS OF THE BOARD OF EDUCATION (SOUTH KENSINGTON) (PAPERBACK)



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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1904 edition. Excerpt: .by it, consequently it is largely used in organic chemistry as a solvent. The commercial alcohol usually contains acetone, from which it may be freed by heating with fused calcium chloride, when the latter combines with the methyl alcohol to form a crystalline compound $\text{CaCl}\cdot\text{CH}_3\text{O}$. This body is stable at 100, so that on heating to this temperature the acetone is driven off. The $\text{CaCl}_2\cdot 4\text{CH}_3\text{O}$ is then treated with water in a retort and distilled. Water and methyl alcohol pass over, and the latter can be obtained in an anhydrous condition by being placed in contact with quicklime, and re-distilled. Methyl alcohol is easily oxidized to formaldehyde and formic acid. 1. When heated in an open dish over a free flame, methyl alcohol volatilizes, takes fire and burns with a pale blue flame, and leaves no residue. 2. To a small quantity of strong sulphuric acid in a test tube, add a little powdered potassium bichromate and finally a small quantity of methyl alcohol. A vigorous action will at once ensue and the solution will be turned green owing to the formation of chromium sulphate. The solution will now contain formic acid. A little formaldehyde may be produced, but the heat of the reaction will be sufficient to drive this off. If the formic acid is present in the solution, the latter must be neutralized and the methyl alcohol distilled off from it, before this test is performed. $2\text{K}_2\text{Cr}_2\text{O}_7 + 8\text{H}_2\text{SO}_4 + 3\text{CH}_3\text{OH} = \text{H}\cdot\text{COOH} + \dots$

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